References and Notes

- (1) A brief presentation of this work has been made: Ortiz de Montellano, P. A brief presentation of this work has been made: Ortiz de Montenatio, P. R.; Vinson, W. A. "Abstracts of Papers", 175th National Meeting of the American Chemical Society, Anaheim, Calif., March 1976; American Chemical Society: Washington, D.C., 1976; ORG 209. "Blochemistry Involving Carbon–Fluorine Bonds", Filler, R. Ed.; American Discriminal Octivity December 2010.
- (2)
- Biochemistry involving Carbon Producting bolids, Finel, R. Ed., American Chemical Society: Washington, D. C., 1976.
 Recent examples follow. (a) Poulter, C. D.; Satterwhite, D. M. *Biochemistry* 1977, *16*, 5470. (b) Gent, M. P. N.; Ho, C. *ibid.* 1978, *17*, 3023. (c) Silver-man, R. B.; Abeles, R. H. *Ibid.* 1977, *16*, 5515. (d) Napoli, J. L.; Fivizzani, M. A.; Schnoes, H. K., DeLuca, H. F. *ibid.* 1978, *17*, 2387. (e) Kollonitsch, H. B. Parceh, J. Am. Chem. Soc. *1026*, *98*, 5591.
- J.; Barash, L. J. Am. Chem. Soc. **1976**, *98*, 5591. Sheppard, W. A.; Sharts, C. M. "Organic Fluorine Chemistry", Benjamin: New York, 1969. (4)
- For example, apart from use of fluorInated isoprenyl pyrophosphates in studies of terpene blosynthesis,^{3a} rapid formation of acyl fluorides from (5) lpha-difluoro alcohols suggests use of their esters as suicidal substrates for
- ca-diffuence and phosphatases.
 (a) Haszeldine, R. N. Nature (London) 1951, 168, 1028. (b) Andreades, S.;
 England, D. C. J. Am. Chem. Soc. 1961, 83, 4670. (c) Mitsch, R. A.; Robertson, J. E. J. Heterocycl. Chem. 1965, 2, 152. (d) Adcock, J. L.; Beh, R. A., Lagow, R. J. J. Org. Chem. 1975, 40, 3271.
 Seppelt, K. Angew. Chem., Int. Ed. Engl. 1977, 16, 322.
 In addition to triffueromethanol. Z partinerovyclobutanal a structure stabilized (6)
- In addition to trifluoromethanol,⁷ perfluorocyclobutanol, a structure stabilized by ring strain and multiple fluorine substitution, has been reported.^{6b} (8) (9)
- Schlosser, M. Tetrahedron 1978, 34, 3. (b) Mathey, F.; Bersoam, J. *ibid.* 1975, 31, 391. (c) Middleton, W. J. J. Org. Chem. 1975, 40, 574. (d) Olah,
 G. A.; Nojima, M.; Kerekes, I. J. Am. Chem. Soc. 1974, 96, 925.
- (10) (a) Brody, E. P., Gutsche, C. D. *Tetrahedron* **1977**, *33*, 723. (b) McCormick, J. P.; Barton, D. L. *ibid*. **1978** *34*, 325.
 (11) Ortiz de Montellano, P. R.; Vinson, W. A., unpublished work.
- (12) Evans, D. A.; Carroll, G. L.; Truesdale, L. K. J. Org. Chem. 1974, 39,
- (13) All structural assignments are consistent with complete spectroscopic and analytical data, including, where applicable, ¹⁹F NMR data.
- and analytical data, including, where applicable, ¹⁹F NMF Naae, D. G.; Burton, D. J. Synth. Commun. **1973**, *3*, 197
- (15) Compound **1a**: ¹H NMR (CDCl₃) 1.43 (d, J = 2 Hz, 3 H, Me at C-3), 1.63 and Compound 1a: A NMA (GLO₃) 1.45 (d, J = 2.47, 5 H, Me at C-3), 1.55 and 1.72 (2 s, 9 H, allylic Me), 2.05 (br m, 8 H, allylic CH₂), 2.15 (s, H, OH), 4.43 (d of d, J = 26, 6 Hz, H, CH==CF₂), 5.20 ppm (m, 2 H, vinyl H); ¹⁹F NMR (CFCl₃ reference) 84.9 (d of d, J = 46, 26 Hz, cls F), 86.3 ppm (d of d, J = 46, 6 Hz, trans F); CIMS m/e 241 (MH⁺ - H₂O). Anal. Calcd for C₁₅H₂₄F₂O: C, 69.73; H, 9.36. Found: C, 69.63; H, 9.12.
- Drakesmith, F. G.; Richardson, R. D.; Stewart, D. J.; Tarrant, P. J. Org. Chem. (16)1968, *33*, 286.
- (a) Babler, J. H.; Olsen, D. O. Tetrahedron Lett. 1974, 351. (b) Babler, J. (17)H.; Coghlan, M. J.; Giacherio, D. J. J. Org. Chem. 1977, 42, 2172.
 Höfle, G.; Steglich, W.; Vorbruggen, H. Angew. Chem., Int. Ed. Engl. 1978,
- 17, 569.
- (19) Liotta, C. L.; Harris, H. P.; McDermott, M.; Gonzales, T.; Smith, K. Tetra-hedron Lett. 1974, 2417.
- Compound **7a**: ¹H (CDCl₃) 1.63 and 1.70 (2 s, 9 H, allylic Me), 1.88 (m, 3 H, Me at C-3), 2.07 (br m, 8 H, allylic CH₂), 2.15 (s, 3 H, MeCO), 5.17 (m, 2 H, vinyl H), 5.60 ppm (t, J = 10 Hz, H, C-2 proton); ¹⁹F NMR (CFCl₃ reference) 63.3 and 63.7 ppm (2 d, J = 10 Hz each, 2*E* and 2*Z* isomers); CIMS *m*/e 241 (MH⁺ → HOAc). Anal. Calcd for C₁₇H₂₆F₂O₂: C, 67.97; H, 8.73. (20) Found: C, 68.36; H, 8.75.
- (21) Vinson, W. A. Ph.D. Thesis, University of California at San Francisco, Sept 1978.
- Compound **5a** (*z*): HINMR (CDCl₃) 1.38 (d of t, J = 1, 7 Hz, 6 H, ethoxy Me), 1.63 and 1.70 (2 s, 9 H, allylic Me), 1.92 (*m*, 3 H, Me at C-3), 2.02 (*m*, 6 H, allylic CH₂), 2.13 (*m*, 2 H, C-4 protons), 4.25 (q, J = 7 Hz, 4 H, OCH₂), 5.55 (*m*, 2 H, Vinyl), 5.58 ppm (t, J = 10 Hz, H, C-2 proton); ¹⁹F NMR (SFCl₃) reference) 55.06 ppm (br m); CIMS *m*/e 395 (MH⁺), 375 (MH⁺ HF). Anal. Calcd for C₁₉H₃₃F₂PO₄: C, 57.85; H, 8.43; P, 7.85. Found: C, 58.11; H, 8.28; P, 7.76. (22) Compound 8a (E): ¹H NMR (CDCl₃) 1.38 (d of t, J = 1, 7 Hz, 6 H, ethoxy Me),

Paul R. Ortiz de Montellano,* Wayne A. Vinson

Department of Pharmaceutical Chemistry School of Pharmacy, University of California San Francisco, California 94143 Received December 12, 1978

Structure Elucidation by High Resolution Mass Spectrometry of a Highly Modified Nucleoside from Mammalian Transfer RNA. $N-[(9-\beta-D-Ribofuranosyl-$ 2-methylthiopurin-6-yl)carbamoyl]threonine

Sir:

The chemical and conformational properties of modified nucleosides in tRNA, particularly those in the anticodon region, are thought to play a role in the biological functions of transfer RNA.^{1,2} As interest in structure and function of eukaryotic tRNA grows, the structure elucidation of new nucleosides becomes more difficult owing to the complexity of

0002-7863/79/1501-2224\$01.00/0





Unfractionated rabbit liver tRNA (5 g) was incubated with nuclease P1 (pH 5.0, 37 °C, 2 h) and the resulting hydrolysate fractionated by DEAE-cellulose (DE-52) with a linear NaCl gradient (0-0.2 M, pH 7.5) in the presence of 7 M urea. Fractions containing the unknown nucleoside N as pNpA were converted into mononucleotides (snake venom phosphodiesterase, pH 7.5, 37 °C, 18 h) which were separated by DEAE Sephadex A-25 (pH 7.8, 0.1-0.7 M, triethylammonium bicarbonate gradient). Nucleotide pN (6 A_{260} units) was purified by two-dimensional paper electrophoresis and chromatography (first run, 30 V/cm for 5 h with 5% acetic acid (adjusted to pH 3.5 by pyridine); second run, isobutyric acid-0.5 M NH₃ (5:3 v/v)) and then dephosphorylated by alkaline phosphomonoesterase.4

Permethylation of the unknown nucleoside using methylsulfinyl carbanion with CD₃I or CH₃I⁵ each yielded two



products which were completely fractionated by vaporization from the mass spectrometer probe at 100 (2, 3) and 150 °C (4, 5), while trimethylsilylation⁶ produced a single derivative $(Si(CH_3)_3, 6; Si(CD_3)_3, 7).$



6, $R^1 = Me_3Si$ (TMS), $R^2 = H$, 818.3184 (-1.2)

© 1979 American Chemical Society

Communications to the Editor

Complete high resolution mass spectra were photographically recorded.⁷ Exact molecular masses of permethyl⁵ and trimethylsilyl⁸ derivatives were established by calculating mean values derived from experimentally measured molecular ion and base series of ions,9 in which the unknown parameter (the exact mass of the base) is associated in each measurement with known mass differences (e.g., $M - CH_3$, base + H). Determination of the number of blocking groups introduced by comparison of molecular weights (2 vs. 3, 6 vs. 7) then permitted calculation of the masses of 1 and its degradation product 8, as shown.

The structure elucidation of nucleoside N followed the characterization of 8 based on the premise that the identity of 8 would permit selection of the basic skeleton from the four major bases. The exact mass of 2 results in 11 computer generated candidates for composition, which can be narrowed by application of a set of restrictions which are generally applicable to nucleosides: (1) total number of rings and double bonds between 4 and 12, (2) oxygen ≥ 4 , (3) nitrogen ≥ 2 , (4) the nitrogen rule. As a result, no candidates emerge which contain only C, H, D, N, and O, while inclusion of S results in a single possibility for 2: $C_{16}H_{10}D_{15}N_5O_4S$. The mass spectrum of 2 shows an unmodified sugar (m/e 149, 183),⁵ an ion of m/e 120 (C₆H₄D₅O₂) generally characteristic of adenosine derivatives,⁵ and loss of CD_2ND from the base + H species, thus revealing $\mathbf{8}$ to have a free exocyclic amino group.^{5,10} These data, plus a detailed analysis of the high resolution mass spectra of 2 and 3, lead to structure 8. An isomer bearing a methylenethiol function is excluded by the extent of methylation (5 instead of 6); substitution of C-8 is excluded by incorporation of one deuterium (D₂O, 100 °C for 1 h, then cold H₂O).

The mass difference between 1 and 8 (146.0467) permits three compositions within the limits $C_{\leqslant 10} H_{\leqslant 20} N_{\leqslant 5} O_{\leqslant 6} S_{\leqslant 1}$ $(\pm 0.004 \text{ amu})$: C₆H₄N₅, C₃H₈N₅S, and C₅H₈NO₄. The N₅



candidates are rejected as structurally implausible. The third corresponds to threonine plus CO, leading to structure 1. Support for the N^6 -carbamoylthreonine structure is gained from the high resolution spectra of 6 and 7, which includes fragment ions of m/e 612, 684, and 701,¹¹ which have analogy in the mass spectra of N-[(9- β -D-ribofuranosylpurin-6-yl)carbamoyl]threonine $(t^6A)^{12}$ and its N⁶-methyladenosine analogue ($mt^{6}A$).¹³



701.2462 (0.2)

The present finding is the first case in which the methylthio group has been found in mammalian tRNA,14,15 while t⁶Å and mt⁶A occur in both prokaryotic and eukaryotic sources.¹⁶ The new nucleoside $(1, ms^2t^6A)$ has recently been located adjacent to the 3' position of the anticodon in tRNA₃Lys from rabbit liver.¹⁷ Its presence thus satisfies the empirical rule that tRNAs

which contain t⁶A or its derivatives recognize messenger RNA codons which begin with A.^{1,2}

Acknowledgment. This work was supported by NATO Grant No. 907 to H.J.G., and NIH Grant CA 18024 to J.A.M.

References and Notes

- S. Nishimura, *Prog. Nucleic Acid Res. Mol. Biol.*, **12**, 49 (1972).
 J. A. McCloskey and S. Nishimura, *Acc. Chem. Res.*, **10**, 403 (1977).
 Exact masses shown are experimental values followed by found minus
- threoretical difference in millimass units. F. Harada, F. Kimura, and S. Nishimura, *Biochemistry*, **10**, 3269 (1971). (5) D. L. von Minden and J. A. McCloskey, J. Am. Chem. Soc., 95, 7480
- (1973).
- S. E. Hattox and J. A. McCloskey, Anal. Chem., 46, 1378 (1974).
- Varian MAT 731 mass spectrometer, 70 eV. J. A. McCloskey, A. M. Lawson, K. Tsuboyama, P. M. Krueger, and R. N. Stillwell, *J. Am. Chem. Soc.*, **90**, 4182 (1968). (8)
- (9) J. A. McCloskey in "Basic Principles in Nucleic Acid Chemistry", Vol. I, P. O. P. Ts'o, Ed., Academic Press, New York, 1976, Chapter 3.
- (10) D. L. von Minden, J. G. Liehr, M. H. Wilson, and J. A. McCloskey, J. Org. Chem., 39, 285 (19).
- (11) Location of the threonyl Me₃Si group in the m/ e 684 ion is uncertain and unimportant, the latter because of its tendency to rearrange during fragmentation.6
- (12) H. Kasai, K. Murao, S. Nishimura, J. G. Liehr, P. F. Crain, and J. A. McCloskey, *Eur. J. Biochem.*, **69**, 435 (1976), and supplemental material A.O. 553: Archives originales du centre de documentation du C.N.R.S., -75971 Paris-Cedex 20, France.
- (13) F. Kimura-Harada, D. L. von Minden, J. A. McCloskey, and S. Nishimura, Biochemistry, 11, 3910 (1972).
- (14) Prokaryotic and plant sources contain 2-methylthio derivatives of N⁶-(3-methyl-2-butenyl)adenosine and N⁶-(4-hydroxy-3-methyl-2-butenyl)adenosine: (a) F. Harada, H. J. Gross, F. Kimura, S. H. Chang, S. Nishimura, and U. L. RajBhandary, *Biochem. Biophys. Res. Commun.*, **33**, 299 (1968); (b) W. J. Burrows, D. J. Armstrong, F. Skoog, S. M. Hecht, J. T. A. Boyle, N. J. Leonard, and J. Occolowitz, *Biochemistry*; **8**, 3071 (1969); (c) D. J. Armstrong, P. K. Evans, W. J. Burrows, F. Skoog, J.- F. Petit, J. L. Dahl, T. Steward, and J. L. Strominger, J. Blol. Chem., 245, 2922 (1970); (d) W. J. Burrows, D. J. Armstrong, M. Kaminek, F. Skoog, R. M. Bock, S. M. Hecht, L. G. Dammann, N. J. Leonard, and J. Occolowitz, *Biochemistry*, 9, 1867 (1970); (e) W. J. Burrows, F. Skoog, and N. J. Leonard, *ibid.*, 10, 2189 (1971); (f) H. J. Vreman, F. Skoog, C. R. Frihart, and N. J. Leonard, *Plant* Physiol., 49, 848 (1972); (g) B. Thimmappaya and J. D. Cherayil, Biochem. Biophys. Res. Commun., 60, 665 (1974).
- (15) The UV spectrum and thin layer chromatographic properties of 1 (ms²t⁶A) suggest that is may be identical with an unknown nucleoside recently isolated from B. subtills tRNALys: Y. Yamada and H. Ishikura, Nucleic Acids Res., 4, 4291 (1977).
- (16) R. H. Hall and D. B. Dunn in "Handbook of Biochemistry and Molecular Biology", 3rd ed., Vol. 1, G. D. Fasman, Ed., CRC Press, Cleveland, Ohio, 1975, Nucleic Acids, p 216.
- (17) M. Raba, K. Limburg, M. Burghagen, H. J. Gross, J. R. Katze, M. Simsek, J. E. Heckman, and U. L. RajBhandary, Eur. J. Biochem., in press.

Z. Yamaizumi, S. Nishimura

Biology Division, National Cancer Center Research Institute, Chuo-ku, Tokyo, Japan

Klaus Limburg, Manfred Raba, Hans J. Gross

Max-Planck-Institut für Biochemie D-8033 Martinsried bei München, West Germany

P. F. Crain, James A. McCloskey*

Departments of Medicinal Chemistry and Biochemistry University of Utah, Salt Lake City, Utah 84112 Received December 19, 1978

Silyl Ketone Chemistry. A New Regiospecific Route to Silvl Enol Ethers

Sir:

The reversible rearrangement of silvl groups from carbon to oxygen in α -silvl alkoxides $(1 \rightleftharpoons 2)^1$ provides an unusual route to potentially useful carbanions. We are investigating several synthetic applications of this rearrangement and report here the development of a new regiospecific silvl enol ether



© 1979 American Chemical Society